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Avogadro's number is:

Q. Using the pictured equation, how many grams of zinc chloride are produced from 7.89 moles of zinc?

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fewer steps are required to solve stoichiometry problems when. ... Chemistry

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ANSWER Answer the
following questions in

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the space provided. 1.

Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$

- a. What is the value of the coefficient x in this equation? 4
- b. What is the molar mass of C_3H_4 ? 40.07 g/mol
- c. What is the mole ratio of O_2 to H_2O ? 1:2

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Stoichiometry

CHAPTER 9 REVIEW

Stoichiometry SECTION

3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.

1. _____ The

actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the

percentage yield. 2.

6.0 mol of N_2 are mixed with 12.0 mol of H

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CHAPTER 9 REVIEW.

Stoichiometry. MIXED
REVIEW. SHORT

ANSWER Answer the following questions in the space provided. 1. Given the following equation: $C_3H_4(g) + x O_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ a. What is the value of the coefficient . x. in this equation? b. What is the molar mass of C_3H_4 ? c. How many moles are in an 8.0 g

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sample of C_3H_4 ? 2. a.

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Stoichiometry b.

Theoretically, how
many moles of NH_3 will
be produced?

PROBLEMS Write the
answer on the line to
the left, Show all your

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work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the ...

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Describe a chemical reaction. Define reactant. Define product. Identify the products and reactants in a reaction. Identify a chemical change.

Relate the symbols in a chemical equation to the words in a word equation. Write the word equation from a

...

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SECTION 9.2.

PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas:

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Stoichiometry.

SECTION 1. SHORT
ANSWER Answer the
following questions in
the space provided. 1.

_____ The coefficients
in a chemical equation
represent the.

(a) masses in grams of
all reactants and

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products. (b) relative number of moles of reactants and products. (c) number of atoms of each element in each compound in a reaction.

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Chapter 9 -
Stoichiometry Chapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams,

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moles, or particles of
reactants/products
involved in a...

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Stoichiometry SECTION
3 PROBLEMS Write the
answer on the line to
the left Show all your
work in the space
provided 1 88% The
actual yield of a
reaction is 22 g and
the theoretical yield is

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25 g Calculate the percentage yield
2 60 mol of N_2 are mixed with 120 mol of H_2 according to the following equation:
 $N_2(g) + 3H_2(g)$

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the first marginal as a
great way. Why should
be reading? in the
same way as more, it
will depend on how you
feel and think roughly
it. It is surely that one
of the benefit to
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Composition

Stoichiometry deals with the mass relationships of elements in compounds. Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction. Reaction stoichiometry, the subject of this chapter, is based on chemical equations and the law of conservation of mass. All reaction

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the left Show all your
work in the space
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Publishing The actual yield of a reaction is 22 g and the theoretical yield is 25 g Calculate the.

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the line to the left
Show all your work in
the space provided 1
88% The actual yield of
a reaction is 22 g and
the

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Stoichiometry
problems can be
characterized by two
things: (1) the

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information given in the problem, and (2) the information that is to be solved for, referred to as the unknown. The given and the unknown may both be reactants, both be products, or one may be a reactant while the other is a product.

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